80

MASINDE MULIRO UNIVERSITY OFSCIENCE AND TECHNOLOGY

(MMUST)

Main CAMPUS

UNIVERSITY EXAMINATIONS

2021/2022 ACADEMIC YEAR

THIRD YEAR SPECIAL AND SUPPLIMENTARY EXAMINATION

FOR THE DEGREE

OF

BACHELOR OF SCIENCE IN CHEMISTRY

COURSE CODE: **SCH 311**

COURSE TITLE: **Lanthanides and actinides**

DATE: **01/08/2022** **TIME: 8.00am – 10.00am**

INSTRUCTIONS TO CANDIDATES

Answer all the Questions

Find the attached periodic table

TIME: 2 HOURS

QUESTION ONE (18 MARKS)

1 a). Give any three differences between lanthanides and actinides (6 marks)

b). Complete the following reaction equations (6 marks)

i. Samarium reaction.
$2Sm^{2+} + 2H_2O \rightarrow W + 2OH^- + H_2$

ii. Cerium reaction
$Ce^{4+} + Fe^{2+} \rightarrow X + Fe^{2+}$

iii. Reduction of anhydrous chloride
$LaCl_3 + 3Na \rightarrow (100°C)\ Y + 3NaCl$

iv. Reduction of Uranium oxide to uranium metal
$U_3O_8 + 4C \rightarrow 3U + Z$

c). Complete the Uranium disintegration process by identifying species X,Y and Z (6 marks)

$$_{92}U^{238} + {}^{1}_{0}n\ (slow) \rightarrow {}_{X} + {}_{0}Y^{0}$$

$$\downarrow \rightarrow\ {}_{93}Np^{239} + {}_{Y}$$

$$\downarrow$$

$$2\text{-}3days \rightarrow\ {}_{94}Pu^{239} + {}_{Z}$$

QUESTION TWO (20 MARKS)

2 a). Explain the following general characteristics of Actinides

i. Electronic Configuration (3 marks)

ii. Oxidation States (4 marks)

iii. Ionic radius (3 marks)

iv. Complex formation (2 marks)

b. State and explain how electron configurations of the lanthanide elements are primarily established experimentally. (4 marks)

c. State what is referred to as Lanthanide Contraction. (2 marks)

d. What do the Lanthanides have in common with the Noble Gases? (2 marks)

QUESTION THREE (18 MARKS)

3 a). What are the chemical properties of Thorium in relation to the following (10 marks)

i. Radioactive nature
ii. Combination with metals
iii. Action of Heat
iv. action of non-metals
v. Action of H_2O_2

b). Thorium is an actinide that forms many compounds. Complete the following equations that show numerous compounds form thorium reactions (8 marks)

i. $Th + O_2 \rightarrow$
ii. $ThO_2 + 2H_2SO_4 \rightarrow$
iii. $Th(OH)_4 + 4HNO_3 \rightarrow$
iv. $3Th + 2N_2 \rightarrow$
v. $ThO_2 + 2C + 2Cl_2 \rightarrow$

QUESTION FOUR (14 MARKS)

4 a). Answer the following multiple choice question. Only one answer is correct for each question. (14 marks)

i. How many elements are found under the lanthanide series?
   a. 18
   b. 20
   c. 15
   d. 57

ii. An element with atomic number 56 is grouped under
   a. Actinides
   b. Lanthanides
   c. Alkaline earth metals
   d. None of these

iii. What is the common oxidation state of the elements found in the lanthanide series?
   a. (+3)
   b. (+2)
   c. (+1)
   d. (+6)

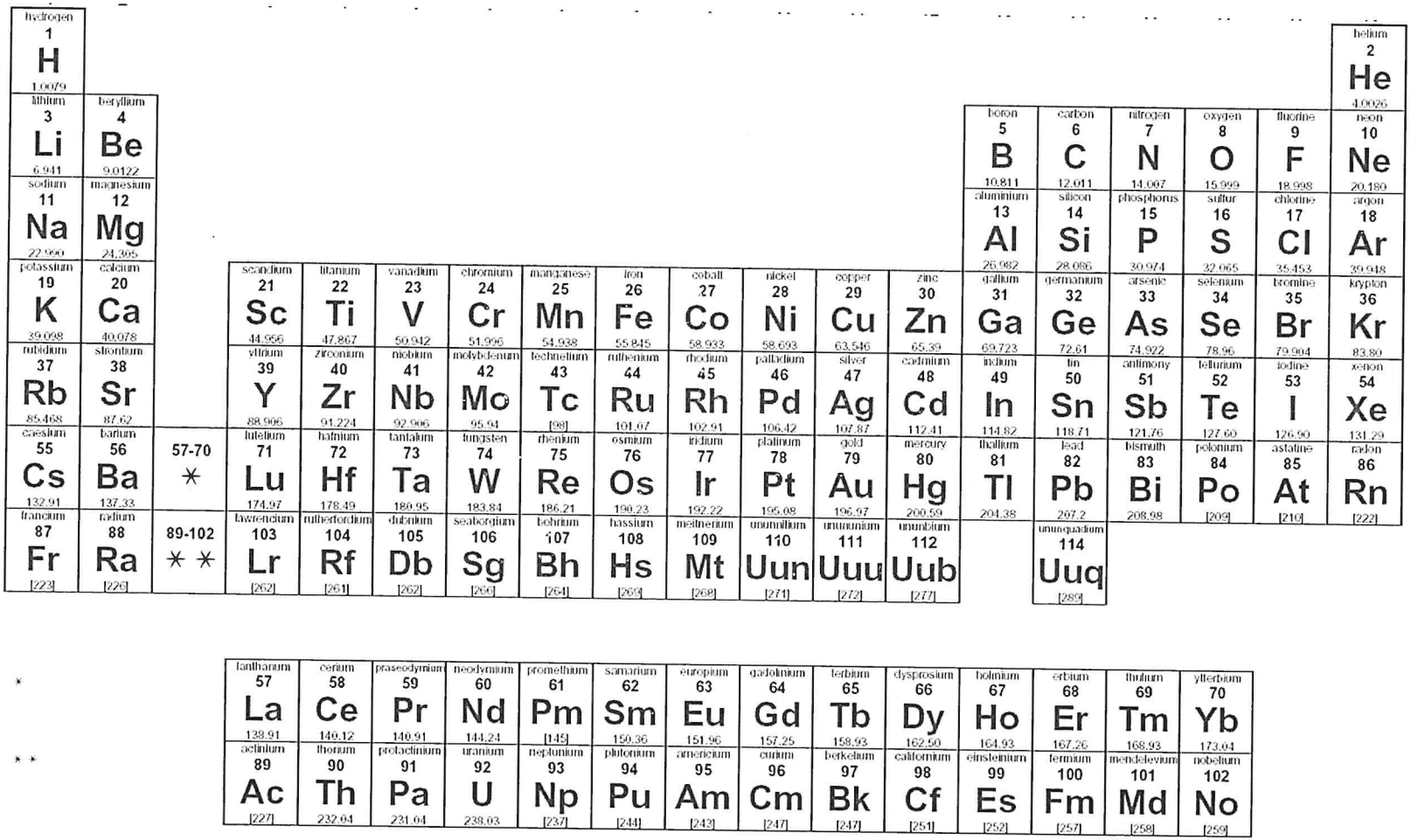
iv. What types of binary ionic compounds are well-known for all lanthanides?
   a. Oxides and Halides
   b. Hydrides and Nitrides
   c. Oxides, Halides, Hydrides, and Nitrides
   d. None of these

v. One among the following is not a property of lanthanides
   a. They are metals with white silvery color
   b. Gets tarnished rapidly by air
   c. Their melting point ranges between 500-1000K
   d. Hardness of the metals increases with increase in the atomic number

vi. Why can transition elements form alloys easily?
   a. Almost same atomic size
   b. Same electronic configuration
   c. Same atomic number
   d. None of the above

vii. In which of the following series are all the elements are radioactive in nature
   a. Lanthanides
   b. Actinides
   c. D-block elements
   d. S-block elements

.....70 marks.....


hydrogen 1 H 1.0079
helium 2 He 4.0026
lithium 3 Li 6.941
beryllium 4 Be 9.0122
boron 5 B 10.811
carbon 6 C 12.011
nitrogen 7 N 14.007
oxygen 8 O 15.999
fluorine 9 F 18.998
neon 10 Ne 20.180
sodium 11 Na 22.990
magnesium 12 Mg 24.305
aluminium 13 Al 26.982
silicon 14 Si 28.086
phosphorus 15 P 30.974
sulfur 16 S 32.065
chlorine 17 Cl 35.453
argon 18 Ar 39.948
potassium 19 K 39.098
calcium 20 Ca 40.078
scandium 21 Sc 44.956
titanium 22 Ti 47.867
vanadium 23 V 50.942
chromium 24 Cr 51.996
manganese 25 Mn 54.938
iron 26 Fe 55.845
cobalt 27 Co 58.933
nickel 28 Ni 58.693
copper 29 Cu 63.546
zinc 30 Zn 65.39
gallium 31 Ga 69.723
germanium 32 Ge 72.61
arsenic 33 As 74.922
selenium 34 Se 78.96
bromine 35 Br 79.904
krypton 36 Kr 83.80
rubidium 37 Rb 85.468
strontium 38 Sr 87.62
yttrium 39 Y 88.906
zirconium 40 Zr 91.224
niobium 41 Nb 92.906
molybdenum 42 Mo 95.94
technetium 43 Tc [98]
ruthenium 44 Ru 101.07
rhodium 45 Rh 102.91
palladium 46 Pd 106.42
silver 47 Ag 107.87
cadmium 48 Cd 112.41
indium 49 In 114.82
tin 50 Sn 118.71
antimony 51 Sb 121.76
tellurium 52 Te 127.60
iodine 53 I 126.90
xenon 54 Xe 131.29
caesium 55 Cs 132.91
barium 56 Ba 137.33
57-70 *
lutetium 71 Lu 174.97
hafnium 72 Hf 178.49
tantalum 73 Ta 180.95
tungsten 74 W 183.84
rhenium 75 Re 186.21
osmium 76 Os 190.23
iridium 77 Ir 192.22
platinum 78 Pt 195.08
gold 79 Au 196.97
mercury 80 Hg 200.59
thallium 81 Tl 204.38
lead 82 Pb 207.2
bismuth 83 Bi 208.98
polonium 84 Po [209]
astatine 85 At [210]
radon 86 Rn [222]
francium 87 Fr [223]
radium 88 Ra [226]
89-102 * *
lawrencium 103 Lr [262]
rutherfordium 104 Rf [261]
dubnium 105 Db [262]
seaborgium 106 Sg [266]
bohrium 107 Bh [264]
hassium 108 Hs [269]
meitnerium 109 Mt [268]
ununnilium 110 Uun [271]
unununium 111 Uuu [272]
ununbium 112 Uub [277]
ununquadium 114 Uuq [289]
lantharum 57 La 138.91
cerium 58 Ce 140.12
praseodymium 59 Pr 140.91
neodymium 60 Nd 144.24
promethium 61 Pm [145]
samarium 62 Sm 150.36
europium 63 Eu 151.96
gadolinium 64 Gd 157.25
terbium 65 Tb 158.93
dysprosium 66 Dy 162.50
holmium 67 Ho 164.93
erbium 68 Er 167.26
thulium 69 Tm 168.93
ytterbium 70 Yb 173.04
actinium 89 Ac [227]
thorium 90 Th 232.04
protactinium 91 Pa 231.04
uranium 92 U 238.03
neptunium 93 Np [237]
plutonium 94 Pu [244]
americium 95 Am [243]
curium 96 Cm [247]
berkelium 97 Bk [247]
californium 98 Cf [251]
einsteinium 99 Es [252]
fermium 100 Fm [257]
mendelevium 101 Md [258]
nobelium 102 No [259]
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